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Any time

A **saturated solution** contains the maximum amount of solute (what is being dissolved) dissolved in an amount of solvent (what you are sticking the solute in). If you add more solute to a **saturated solution**, it will not dissolve. An **unsaturated solution** is opposite of this. It can dissolve more solute.

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Differences Between Each **Solution-Saturated** vs **Unsaturated-Saturated** is when there is a little bit of solute left that hasn't dissolved completely in the solvent.

Unsaturated can be if there were more solute to put in an **unsaturated** solvent it would absorb more solute. **Saturated** is when there is a little bit of solute left that hasn't dissolved

What is the difference between a saturated solution &

<https://www.quora.com/What-is-the-difference-between-a-saturated...>

A **saturated solution** basically implies that all the solvent has been dissolved with the solute, such that if more solute is added, it would not dissolve; this is called a **saturated solution**. An **unsaturated solution** basically means that if you add more solute, it can and will still dissolve in the solvent until it reaches saturation.

Saturated and Supersaturated Solutions - Chemistry |

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We have a **saturated solution**. If we now heat the mixture to 50 °C, the remaining 9 g of glucose will dissolve. At the new temperature, the solubility limit in 100 mL of water is 244 g glucose. With only 100 g of glucose dissolved, the **solution** is now **unsaturated**.

Unsaturated vs Saturated vs Supersaturated solutions ...

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Dec 21, 2015 · An **unsaturated solution** is one in which a little amount of solute has been added to the solvent. A **solution** is said to be **saturated** when a solute is not able to dissolve in the solvent. A supersaturated **solution**, on the other hand, is when the excess of solute is dissolved in the solvent as a result of changes in temperature,

1/10

(1)

Types of Solutions: Saturated, Supersaturated, or Unsaturated

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Types of **Solutions: Saturated, Supersaturated, or Unsaturated** . Resource ID: CM5L3. Grade Range: 9-12. Sections. **Unsaturated, Saturated, and Supersaturated Examples** . How to Read a Solubility Graph . Practice Reading a Solubility Graph"Part 1 . Practice Reading a Solubility Graph"Part 2 . **Unsaturated, Saturated, and Supersaturated**

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An **unsaturated solution** is when the solute CANNOT dissolve into the solvent. For example, if you have a glass of water, and you pour something like vegetable oil into it, it will not dissolve. So that makes the **solution unsaturated**

solutions tutorial- unsaturated, saturated &

<https://www.youtube.com/watch?v=JQRiIS9INdg>

Apr 03, 2014 · This video screencast was created with Doceri on an iPad. Doceri is free in the iTunes app store. Learn more at <http://www.doceri.com>

Unsaturated Solution: Definition & Examples - Study.com

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Saturated Solution: Definition & Examples ... And it's important to know that in any **solution**, there is usually more solvent than solute; this makes the **solution unsaturated**. Now let's turn our attention to how a solute dissolves in a solvent. This has to do with polarity. Remember that a polar molecule has unequal sharing of electrons.

Examples of Saturated Solution - YourDictionary

[examples.yourdictionary.com/examples-of-saturated-solution.html](https://www.yourdictionary.com/examples-of-saturated-solution.html)

The term **saturated solution** is used in chemistry to define a **solution** in which no more solvent can be dissolved. See some examples of **saturated solution** to understand what a **saturated solution** is.

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What is a supersaturated solution?



What is supersaturated in chemistry?



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